

## The Nature of the Active Sites on $Ni/CeO_2$ Catalysts for Methane Conversions: When Theory and Experiment go Hand in Hand

M. VERÓNICA GANDUGLIA-PIROVANO<sup>a</sup>, PABLO LUSTEMBERG<sup>a, b</sup> <sup>a</sup> Institute of Catalysis and Petrochemistry, CSIC, Madrid, Spain.

<sup>b</sup> Institute of Physics Rosario, CONICET-UNR, Rosario, Santa Fe, Argentina.

email: vgp@icp.csic.es

Methane (CH<sub>4</sub>) is difficult to convert to upgraded fuels or chemicals because of the strength of the C-H bonds. Its activation and reforming with CO<sub>2</sub> to syngas or with H<sub>2</sub>O to produce hydrogen, as well as its direct oxidation to methanol, is increasingly important because CH<sub>4</sub> is a potent greenhouse gas. Ni at low loadings on CeO<sub>2</sub> are active for CH<sub>4</sub> activation at room-temperature and for these methane conversion reactions [1-3]. Revealing the nature of the active sites in such systems is paramount to a rational design of improved catalysts. Here we show that the most active sites are Ni atoms in clusters at step edges on the CeO<sub>2</sub> surface, with a small size. We show that the size and morphology of the supported nanoparticles together with strong Ni-support bonding and charge transfer at the step edge are key to the high catalytic activity towards methane activation, allowing much better activity than expected from earlier linear scaling relations [4-6]. The emphasis is here put on theoretical studies in combination with experiments using ambient pressure X-ray photoelectron spectroscopy, adsorption calorimetry and surface analysis measurements.

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